Alkalinity Unknown Data Sheet

Vc	blunteer (School) Group:	Test Date:/	_/			
All	kalinity Unknown Batch:		-			
T	est 1:					
	art 1 – Phenolphthalein Alkalinity Did you clean your Erlenmeyer flask and cylinder with deionized wate	r? Yes	No			
2.	Amount of Alkalinity Unknown used:		mL			
3.	Add 15 drops of phenolphthalien indicator.					
4.	Did the sample turn pink after phenolphthalien was added to flask?	Yes	No			
	YES \rightarrow continue on to step 5. NO \rightarrow record phenolphthalein alkalinity as 0.0 mg/L or	n line 5, and then g	o to part 2.			
5.	Titrate from a pink to a clear. Read your burette and record mL of H2	SO4 you added.	mL			
6.	Multiply mL of H2SO4 used by 40. Record this as the phenolphthalein alkalinity:					
	Phenolphthalein Alkalinity	Result	_mg/L CaCo₃			
To	DO NOT re-zero your burette between phenolphthalein otal is a combination of them both!!	alkalinity and total	<u>alkalinity!</u>			
Pa	<u>rrt 2 – Total Alkalinity</u>					
7.	Add 6 drops of BGMR solution. Did solution turn blue/turquoise?		No			
8.	Titrate from turquoise to pink-gray, record the mL of H2SO4 added.		ml H2SO4			
	Read burette carefully! This is the end point of titration.					
9.	Multiply mL of H2SO4 (line 8) used by 20. (mL H2SO4 use	d) x 20 =	mg/L CaCO3			
Co	omments					
Vo	olunteer Signature:	Date recorded				
Tu	Irn page over for test 2.					

Alkalinity Unknown Batch: _____

Test 2:

Ikalinity Unknown True Val	ue (to be completed by	RW Staff).	mg/	/L CaCO3
est 1 result:	Test 2 result:	Avera	ıge:	
verage the 2 test results:				
Volunteer Signature:			rded	
Comments				
9. Multiply mL of H2SO4	(line 8) used by 20.	(mL H2SO4 used) x 20 =		_mg/L CaCO3
Read burette care	efully! This is the end p	point of titration.		
8. Titrate from turquoise t	Titrate from turquoise to pink-gray, record the mL of H2SO4 added.			_ml H2SO4
7. Add 6 drops of BGMR	solution. Did solution tur	n blue/turquoise?	Yes	No
Part 2 – Total Alkalir	nity			
<u>!! DO NOT re-zero yo</u> Total is a combinatio		phenolphthalein alkalini	ty and total a	<u>lkalinity!</u>
	Phenolpl	hthalein Alkalinity Result	mg	l/L CaCo₃
6. Multiply mL of H2SO4	used by 40. Record this	as the phenolphthalein alkalinity	:	
5. Titrate from a pink to a	clear. Read your burette	e and record mL of H2SO4 you a	added	mL
If YES \rightarrow continue or If NO \rightarrow record phe		nity as 0.0 mg/L on line 5,	and then go	to part 2.
4. Did the sample turn pir	nk after phenolphthalien	was added to flask?	Yes	No
3. Add 15 drops of pheno	Iphthalien indicator.			
2. Amount of Alkalinity Ur	nknown used:			mL